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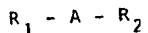
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(54) A fluorine-containing compound and a method of preparing the same.

(57) A novel fluorine-containing compound of the formula



wherein A is  $\{CF_2CFX\}_m$ ,  $\{CF_2CF\}_m$ ,  $\{CF_2CF\}_m$  or  $\{CF_2CF=CFCF_2\}_m$ ;

$$\begin{array}{c} | \\ CF_3 \\ | \\ CF \\ || \\ CF_2 \end{array}$$

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R<sub>1</sub> and R<sub>2</sub> are each hydrogen or R<sub>3</sub>OOG(CH<sub>2</sub>)<sub>n</sub>, while R<sub>1</sub> and R<sub>2</sub> may not be hydrogen at the same time; X is hydrogen or halogen; R<sub>3</sub> is hydrogen or a lower alkyl group; m is an integer of 1 to 5; and n is an integer of 4 to 11, and a process for preparing the same.

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A FLUORINE-CONTAINING COMPOUND AND  
A METHOD OF PREPARING THE SAME

This invention relates to fluorine-containing compounds and more particularly to fluorine-containing carboxylic acids and esters thereof and to a process 5 for preparing the same.

The fluorine-containing compounds of this invention are those represented by the formula



10 wherein A is  $\{CF_2CFX\}_m$ ,  $\{CF_2CF\}_m$ ,  $\{CF_2CF\}_m$  or  
 $\begin{array}{c} CF_3 \\ | \\ CF \\ || \\ CF_2 \end{array}$

$\{CF_2CF=CFCF_2\}_m$ ;  $R_1$  and  $R_2$  are each hydrogen or  $R_3OOC(CH_2)_n$ , while  $R_1$  and  $R_2$  may not be hydrogen at the same time; 15 X is hydrogen or halogen;  $R_3$  is hydrogen or a lower alkyl group; m is an integer of 1 to 5; and n is an integer of 4 to 11.

The fluorine-containing compounds of this invention and esters thereof are novel compounds which we have developed for the first time and which are not described in any literature.

20 The compounds prepared by oxidizing a telomer of tetrafluoroethylene and methanol are known as fluorine-containing carboxylic acids. These conventional compounds have carboxyl groups linked directly with fluorine-

containing groups. Such known carboxylic acids or esters thereof are chemically unstable in a considerable degree. Further the esters thereof are susceptible to hydrolysis and these disadvantages seriously restrict  
5 the use of the prior art esters thereof. On the other hand, the compounds of this invention have long-chain methylene groups positioned between the carboxyl groups and the fluorine-containing functional groups and thus are essentially different in chemical structure from  
10 the conventional fluorine-containing carboxylic acid compounds. No compound has been known heretofore in which long-chain methylene groups are present between fluorine-containing groups and carboxyl groups. The novel compounds of this invention are remarkably  
15 superior in chemical stabilities to conventional carboxylic acid compounds. Moreover, the esters of this invention are conspicuously stable against hydrolysis and therefore find wide applications, because stable carboxyl groups can be utilized.

20 The novel fluorine-containing compounds of this invention have remarkable emulsifying properties and can be used singly as an emulsifier. The present compounds are also usable as an intermediate for the preparation of synthetic resins. The synthetic resins  
25 thus prepared have excellent emulsifying properties

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and thus are well suitable for use in the preparation of emulsion coating compositions.

Table 1 below gives typical examples of the fluorine-containing carboxylic acids of this invention  
5 and their esters.

Table 1

1.	$\text{CH}_3\text{OC}(\text{CH}_2)_5\text{CF}_2\text{CF}-\text{H}$    0 CF <sub>3</sub>	(7-di-8-monofluoro)-9-trifluoro- pelargonic acid methyl
2.	$\text{CH}_3\text{OC}(\text{CH}_2)_5\text{CF}_2\text{CF}-(\text{CH}_2)_5\text{COCH}_3$    0 CF <sub>3</sub>	(7-di-8-monofluoro)-8-trifluoromethyl- tetradecadiacid dimethyl
3.	$\text{HO}(\text{CH}_2)_5\text{CF}_2\text{CF}-\text{H}$    0 CF <sub>3</sub>	(7-di-8-monofluoro)-9-trifluoro- pelargonic acid
4.	$\text{HO}(\text{CH}_2)_5\text{CF}_2\text{CF}-(\text{CH}_2)_5\text{COH}$    0 CF <sub>3</sub>	(7-di-8-monofluoro)-8-trifluoromethyl- tetradecadiacid
5.	$\text{CH}_3\text{OC}(\text{CH}_2)_5\text{CF}_2\text{CHF}-\text{H}$    0	7-di-8-monofluorooctanoic acid methyl
6.	$\text{CH}_3\text{OC}(\text{CH}_2)_5\text{CF}_2\text{CHF}-\text{H}$    0	7,9-di-8,10-monofluorodecanoic acid methyl
7.	$\text{CH}_3\text{OC}(\text{CH}_2)_5\text{CF}_2\text{CHF}-\text{H}$    0	7,9,11-di-8,10,12-monofluorododecanoic acid methyl
8.	$\text{CH}_3\text{OC}(\text{CH}_2)_5\text{CF}_2\text{CHF}-(\text{CH}_2)_5\text{COCH}_3$    0	7-di-8-monofluorotetradecadiacid dimethyl

9.	$\text{CH}_3\text{OC}-(\text{CH}_2)_5-(\text{CF}_2\text{CHF})_2-(\text{CH}_2)_5-\text{COCH}_3$	7,9-di-8,10-monofluorohexadecadiacid
10.	$\text{CH}_3\text{OC}-(\text{CH}_2)_5-(\text{CF}_2\text{CHF})_3-(\text{CH}_2)_5-\text{COCH}_3$	7,9,11-di-8,10,12-monofluoro-octadecadiacid dimethyl
11.	$\text{HO}-(\text{CH}_2)_5-\text{CF}_2\text{CHF}-\text{H}$	7-di-8-monofluoroctanoic acid
12.	$\text{HO}-(\text{CH}_2)_5-(\text{CF}_2\text{CHF})_2-\text{H}$	7,9-di-8,10-monofluorodecanoic acid
13.	$\text{HO}-(\text{CH}_2)_5-(\text{CF}_2\text{CHF})_3-\text{H}$	7,9,11-di-8,10,12-monofluoro-dodecanoic acid
14.	$\text{HO}-(\text{CH}_2)_5-\text{CF}_2\text{CHF}-(\text{CH}_2)_5-\text{COH}$	7-di-8-monofluorotetradecadiacid
15.	$\text{HO}-(\text{CH}_2)_5-(\text{CF}_2\text{CHF})_2-(\text{CH}_2)_5-\text{COH}$	7,9-di-8,10-monofluorohexadecadiacid
16.	$\text{HO}-(\text{CH}_2)_5-(\text{CF}_2\text{CHF})_3-(\text{CH}_2)_5-\text{COH}$	7,9,11-di-8,10,12-monofluoro-octadecadiacid

17.  $\text{CH}_3\text{OC}-(\text{CH}_2)_5-\text{CF}_2\text{CF}_2-\text{H}$  7,8-difluoroctanoic acid methyl

18.  $\text{CH}_3\text{OC}-(\text{CH}_2)_5-(\text{CF}_2\text{CF}_2)_2-\text{H}$  7,8,9,10-difluorodecanoic acid methyl

19.  $\text{CH}_3\text{OC}-(\text{CH}_2)_5-(\text{CF}_2\text{CF}_2)_3-\text{H}$ , 7,8,9,10,11,12-difluorododecanoic acid methyl

20.  $\text{CH}_3\text{OC}-(\text{CH}_2)_5-\text{CF}_2\text{CF}_2-(\text{CH}_2)_5-\text{COCH}_3$  7,8-difluorotetradecadiacid dimethyl

21.  $\text{CH}_3\text{OC}-(\text{CH}_2)_5-(\text{CF}_2\text{CF}_2)_2-(\text{CH}_2)_5-\text{COCH}_3$  7,8,9,10-difluorohexadecadiacid dimethyl

22.  $\text{CH}_3\text{OC}-(\text{CH}_2)_5-(\text{CF}_2\text{CF}_2)_3-(\text{CH}_2)_5-\text{COCH}_3$  7,8,9,10,11,12-difluoroctadecadiacid dimethyl

23.  $\text{HO}-(\text{CH}_2)_5-\text{CF}_2\text{CF}_2-\text{H}$  7,8-difluoroctanoic acid

24.  $\text{HO}-(\text{CH}_2)_5-(\text{CF}_2\text{CF}_2)_2-\text{H}$  7,8,9,10-difluorodecanoic acid

7,8,9,10,11,12-difluorododecanoic acid

7,8-difluorotetradecadiacid

7,8,9,10-difluorohexadecadiacid

7,8,9,10,11,12-difluoroctadecadiacid

7,7,8-trifluoro-8-chlorooctanoic acid methyl

7,7,8,9,10-hexafluoro-8,10-dichlorododecanoic acid methyl

7,7,8,9,10,11,12-nonylfluoro-8,10,12-trichlorododecanoic acid methyl

7,7,8-trifluoro-8-chlorotetradecadiacid dimethyl

25.  $\text{HOC} - (\text{CH}_2)_5 - (\text{CF}_2\text{CF}_2)_3 - \text{H}$

26.  $\text{HOC} - (\text{CH}_2)_5 - \text{CF}_2\text{CF}_2 - (\text{CH}_2)_5 - \text{COH}$

27.  $\text{HOC} - (\text{CH}_2)_5 - (\text{CF}_2\text{CF}_2)_2 - (\text{CH}_2)_5 - \text{COH}$

28.  $\text{HOC} - (\text{CH}_2)_5 - (\text{CF}_2\text{CF}_2)_3 - (\text{CH}_2)_5 - \text{COH}$

29.  $\text{CH}_3\text{OC} - (\text{CH}_2)_5 - \text{CF}_2\text{CFClH}$

30.  $\text{CH}_3\text{OC} - (\text{CH}_2)_5 - (\text{CF}_2\text{CFCl})_2\text{H}$

31.  $\text{CH}_3\text{OC} - (\text{CH}_2)_5 - (\text{CF}_2\text{CFCl})_3 - \text{H}$

32.  $\text{CH}_3\text{OC} - (\text{CH}_2)_5 - \text{CF}_2\text{CFCl}(\text{CH}_2)_5 - \text{COCH}_3$

33.  $\text{CH}_3\text{OC}-(\text{CH}_2)\overset{\parallel}{\underset{\text{O}}{|}}\text{OC}-(\text{CF}_2\text{CFC1})_2-(\text{CH}_2)\overset{\parallel}{\underset{\text{O}}{|}}\text{COCH}_3$  7,7,8,9,9,10-hexafluoro-8,10-dichlorotetradecadiacid dimethyl

34.  $\text{CH}_3\text{OC}-(\text{CH}_2)\overset{\parallel}{\underset{\text{O}}{|}}\text{OC}-(\text{CF}_2\text{CFC1})_3-(\text{CH}_2)\overset{\parallel}{\underset{\text{O}}{|}}\text{COCH}_3$  7,7,8,9,9,10,11,11,12-nonylfluoro-8,10,12-trifluorotetradecadiacid dimethyl

35.  $\text{HOC}-(\text{CH}_2)\overset{\parallel}{\underset{\text{O}}{|}}\text{CF}-\text{CF}_2\text{CFC1}-\text{H}$  7,7,8-trifluoro-8-chlorooctanoic acid

36.  $\text{HOC}-(\text{CH}_2)\overset{\parallel}{\underset{\text{O}}{|}}\text{CF}-(\text{CF}_2\text{CFC1})_2-\text{H}$  7,7,8,9,9,10-hexafluoro-8,10-dichlorodecanoic acid

37.  $\text{HOC}-(\text{CH}_2)\overset{\parallel}{\underset{\text{O}}{|}}\text{CF}-(\text{CF}_2\text{CFC1})_3-\text{H}$  7,7,8,9,9,10,11,11,12-nonylfluoro-8,10,12-trichlorododecanoic acid

38.  $\text{HOC}-(\text{CH}_2)\overset{\parallel}{\underset{\text{O}}{|}}\text{CF}-(\text{CF}_2\text{CFC1})-(\text{CH}_2)\overset{\parallel}{\underset{\text{O}}{|}}\text{COH}$  7,7,8-trifluoro-8-chlorotetradecadiacid

39.  $\text{HOC}-(\text{CH}_2)\overset{\parallel}{\underset{\text{O}}{|}}\text{CF}-(\text{CF}_2\text{CFC1})_2-(\text{CH}_2)\overset{\parallel}{\underset{\text{O}}{|}}\text{COH}$  7,7,8,9,9,10-hexafluoro-8,10-dichlorotetradecadiacid

40.  $\text{HOC}-(\text{CH}_2)\overset{\parallel}{\underset{\text{O}}{|}}\text{CF}-(\text{CF}_2\text{CFC1})_3-(\text{CH}_2)\overset{\parallel}{\underset{\text{O}}{|}}\text{COH}$  7,7,8,9,9,10,11,11,12-nonylfluoro-8,10,12-trifluorotetradecadiacid

41.  $\text{CH}_3\text{OC}-(\text{CH}_2)_5-\text{CF}_2\text{CHF}-\text{H}$   
0  
7-di-8-monofluoroctanoic acid methyl

42.  $\text{CH}_3\text{OC}-(\text{CH}_2)_5-(\text{CF}_2\text{CHF})_2\text{H}$   
0  
7,9-di-8,10-monofluorodecanoic acid methyl

43.  $\text{CH}_3\text{OC}-(\text{CH}_2)_5-(\text{CF}_2\text{CHF})_3\text{H}$   
0  
7,9,11-di-8,10,12-monofluorododecanoic acid methyl

44.  $\text{CH}_3\text{OC}-(\text{CH}_2)_5-\text{CF}_2\text{CHF}-(\text{CH}_2)_5-\text{COCH}_3$   
0  
7-di-8-monofluorotetradecadiacid dimethyl

45.  $\text{CH}_3\text{OC}-(\text{CH}_2)_5-(\text{CF}_2\text{CHF})_2-(\text{CH}_2)_5-\text{COCH}_3$   
0  
7,9-di-8,10-monofluorohexadecadiacid dimethyl

46.  $\text{CH}_3\text{OC}-(\text{CH}_2)_5-(\text{CF}_2\text{CHF})_3-(\text{CH}_2)_5-\text{COCH}_3$   
0  
7,9,11-di-8,10,12-monofluoroctadecadiacid dimethyl

47.  $\text{HOOC}-(\text{CH}_2)_5-\text{CF}_2\text{CHF}-\text{H}$   
0  
7-di-8-monofluoroctanoic acid

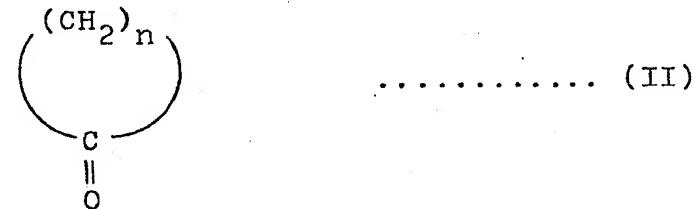
48.  $\text{HOOC}-(\text{CH}_2)_5-(\text{CF}_2\text{CHF})_2\text{H}$   
0  
7,9-di-8,10-monofluorodecanoic acid



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The novel fluorine-containing carboxylic acids of this invention and their esters and prepared by a process comprising the steps of reacting a hydrogen peroxide with a cyclic ketone of the formula

5



(wherein  $n$  is an integer of 4 to 11) to yield a cyclic ketone peroxide; and reacting the cyclic ketone peroxide thus prepared with a fluorine-containing olefin of the formula

10



wherein  $R_4$  is hydrogen, halogen, a trifluoromethyl group or the group  $\text{CF}_2=\text{CF}-$  in the presence of a ferrous salt to prepare a fluorine-containing carboxylic acid of the formula

15



wherein  $A$  is  $\langle \text{CF}_2\text{CFX} \rangle_m$ ,  $\langle \text{CF}_2\text{CF} \rangle_m$ ,  $\langle \text{CF}_2\text{CF} \rangle_m$  or



$\langle \text{CF}_2\text{CF}=\text{CFCF}_2 \rangle_m$ ;  $\text{R}_1$  and  $\text{R}_2$  are each hydrogen or  $\text{R}_3\text{OOC}(\text{CH}_2)_n$ , while  $\text{R}_1$  and  $\text{R}_2$  may not be hydrogen at the same time;  $X$  is hydrogen or halogen;  $\text{R}_3$  is hydrogen or a lower alkyl group;  $m$  is an integer of 1 to 5; and  $n$  is an integer of

4 to 11. Preferably the compounds of the formula (I) are obtained in which the lower alkyl group contains 1 to 4 carbon atoms.

The cyclic ketone of the formula (II) and 5 hydrogen peroxide are reacted in an appropriate ratio which is not particularly limited but widely variable. Usually the latter is used in an amount of 0.5 to 2 moles, preferably about 1 mole, per mole of the former. Exemplary of useful cyclic ketones of the formula (II) 10 are cyclopentanone, cyclohexanone, cycloheptanone, cyclooctanone, cyclononanone, cyclodecanone, cyclo-undecanone, cyclododecanone, etc. The hydrogen peroxide is used usually in the form of an aqueous solution 15 having a concentration of about 30 to about 70%. This reaction (hereinafter referred to as "first reaction") is carried out preferably in the presence of a solvent. Typical of useful solvents are methanol, ethanol, iso-propanol, n-butanol, tert-butanol or like lower alcohols, acetonitrile, benzene, dimethylformamide and the like. 20 Among them, methanol and particularly absolute methanol are preferred. An acid catalyst may be incorporated in the aforesaid reaction mixture. The addition of the acid catalyst gives a higher yield of cyclic ketone peroxide. Examples of useful acid catalysts are sulfuric acid, hydrochloric acid, phosphoric acid and like mineral 25

acids among which a sulfuric acid is desirable. The acid catalyst is used preferably in an amount of about 0.05 to about 0.5 mole per mole of the cyclic ketone of the formula (II). The first reaction is conducted 5 generally at about -70 to about 50°C, preferably at about -10 to about 10°C, and completed usually in about 5 to about 60 minutes.

According to this invention, the cyclic ketone peroxide may be separated from the reaction mixture thus 10 obtained and used in the subsequent reaction (hereinafter referred to as "second reaction"). Alternatively the reaction mixture per se containing the cyclic ketone peroxide therein may be subjected, as it is or as concentrated, to the second reaction. While the ratio of 15 the cyclic ketone peroxide to the fluorine-containing olefin of the formula (III) to be used in the second reaction is not restricted to a narrow range, the latter is employed in a quantity of usually about 1 to about 8 moles, preferably about 2 to 4 moles, per mole of the 20 former. Useful fluorine-containing olefins of the formula (III) are illustrated by trifluoroethene, tetrafluoroethene, 1-chloro-1,2,2-trifluoroethene, hexafluoropropene, hexafluorobutadiene and the like. It is desired to carry out the second reaction in the 25 presence of a suitable solvent. The solvents useful

in the first reaction are usable as the solvent in the second reaction. The second reaction is effected in the presence of a ferrous salt serving as a catalyst. Examples of useful catalysts are ferrous sulfate, 5 ferrous chloride, ammonium ferrous sulfate and the like. The ferrous salt is used in an amount of about 0.1 to about 2 moles, preferably about 0.8 to about 1.2 moles, per mole of the cyclic ketone peroxide. A preferred mode of the reaction is as follows. A specified quantity 10 of the fluorine-containing olefin of the formula (III) as mixed with the solvent is stirred at a temperature below the boiling point of the olefin, generally -70 to 50°C, preferably -40 to -25°C to obtain a suspension. Then a specific amount of ferrous salt is added to the 15 suspension, and the cyclic ketone peroxide is dropwise added to the mixture with stirring. This reaction, if conducted under nitrogen atmosphere, will produce the contemplated product in a high yield. The first reaction is completed generally in about 1 to about 7 20 hours.

With this invention, the cyclic ketone peroxide produced by the reaction of a cyclic ketone with a hydrogen peroxide in lower alcohol is a lower alkoxy cyclic ketone peroxide. Thus the second reaction between 25 the same and the olefin of the formula (III) gives the

present compound of the formula (I) mainly in the form of an ester (namely a compound of the formula (I) wherein R<sub>3</sub> is a lower alkyl group). Among the compounds of the formula (I), carboxylic acid compounds (i.e. those of the formula (I) wherein R<sub>3</sub> is hydrogen) are synthesized by hydrolyzing the esters in the conventional manner.

When reacting the cyclic ketone with the hydrogen peroxide in a solvent other than the lower alcohol, this produces a cyclic ketone peroxide but not an ester thereof. Therefore, on subsequent reaction between the same and the olefin of the formula (III), a compound of the formula (I) is obtained in the form of a carboxylic acid (monobasic acid or dibasic acid). It follows that an ester of the formula (I) can be synthesized by the esterification of the foregoing carboxylic acid in the usual manner. With this invention, the ester of the compound of the formula (I) can be conventionally converted into another ester by ester interchange; for example, methyl esters are transformable into ethyl esters.

The compounds of this invention thus prepared are easily isolated and purified for instance by the extraction with a solvent or distillation or like conventional operation. An excess of olefin of the formula (III), solvent or the like present in the reaction

product can be recovered by a simple distilling procedure and effectively reused.

This invention will be described below in greater detail with reference to Examples to which the 5 invention is not limited.

Example 1

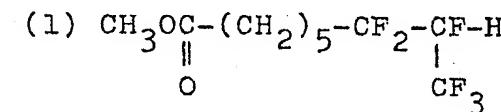
A 100 ml quantity of absolute methanol is placed in a reactor cooled to a temperature of 0 to 5°C and equipped with a stirrer. Thereto are added 14.7 g of cyclohexanone, 1.5 g of concentrated sulfuric acid and 17.0 g of a 30% aqueous solution of hydrogen peroxide. Then the mixture is stirred for 40 minutes to yield methoxycyclohexanone peroxide. A 150 ml quantity of absolute methanol is placed in another reactor equipped with a stirrer and filled with a nitrogen gas, and is cooled to a temperature of -30°C. Thereto are added 46 g of ferrous sulfate heptahydrate ( $FeSO_4 \cdot 7H_2O$ ) and 60 ml of liquefied hexafluoropropene to prepare a suspension. The methoxycyclohexanone peroxide solution is dropwise added to the suspension with stirring while maintaining the temperature of the suspension at -30 + 2°C. After the addition thereof, the mixture is stirred for 2.5 hours and the resulting mixture is left to stand until it is brought to room temperature. The unreacted 20 hexafluoropropene (about 45 ml) is recovered by distillation. 25

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The reaction mixture is separated into an organic layer (upper layer) and an aqueous layer (lower layer). The organic layer (30.6 g) is washed with acid and water, and dried. A fraction (light yellow to colorless) is obtained by distillation at 50 to 200°C/1.4 mmHg.

The fraction is a mixture of long-chain monobasic acid, long-chain ester of monobasic acid, long-chain dibasic acid and long-chain ester of dibasic acid all of which contain fluorine. The mixture is separated into individual compounds by rectification.

Analyses by chromatography, mass spectrometry, etc. confirm the production of the following compounds.



Elementary analysis (for  $\text{C}_{10}\text{H}_{14}\text{O}_2\text{F}_6$ )

15

	C	H	O	F
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Calcd. (%)	42.86	5.00	11.43	40.71
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Found (%)	43.01	5.30	11.19	40.50
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IR:  $1020 \text{ cm}^{-1}$  (C-F),

$1100 \text{ cm}^{-1} - 1230 \text{ cm}^{-1}$  (C-F),

20

$1740 \text{ cm}^{-1}$  (C=O)

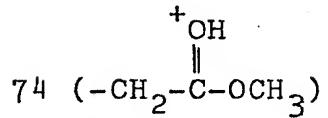
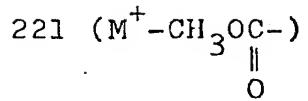
$^1\text{H NMR (CCl}_4\text{): } 3.62\tau$  (s,  $\text{CH}_3\text{O}$ )

$0.90 - 2.60\tau$  ( $\text{CH}_2$  and CFH)

Mass : m/e = 280 ( $\text{M}^+$ )

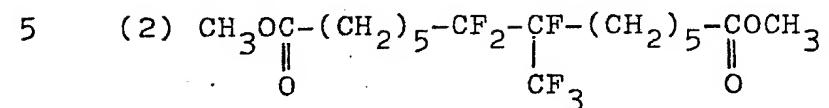
249 ( $\text{M}^+ - \text{CH}_3\text{O}$ )

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Boiling point: 73°C/1.4 mmHg

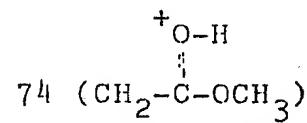
Yield: 17.3 g



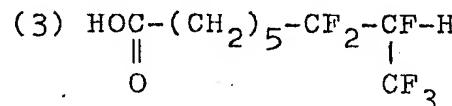
## Elementary analysis

		C	H	O	F
	Calcd. (%)	50.00	6.37	15.69	27.94
10	Found (%)	50.27	6.30	15.69	27.74
	IR:	1020 $\text{cm}^{-1}$		(C-F)	
		1100 - 1230 $\text{cm}^{-1}$		(C-F)	
		1740 $\text{cm}^{-1}$		(C=O)	
15	$^1\text{H}$ NMR ( $\text{CCl}_4$ )	1.2 - 2.1 $\tau$		(bt, $\text{CH}_2$ and CFH)	
		2.23 $\tau$		(t, $\text{CH}_2\text{CO}$ , $J=6.5$ Hz)	
		3.55 $\tau$		(s, $\text{CH}_3\text{O}$ )	
	$^{19}\text{F}$ NMR ( $\text{CCl}_4$ ):			( $\text{CF}_3\text{COOH}$ )	
				-3.9 $\tau$ (t.d, $\text{CF}_3$ , $J=231$ Hz, 14.7 Hz)	
				30.7 $\tau$ (q.d, $\text{CF}_2$ , $J=23.7$ Hz, 6.8 Hz)	
20		101.4 $\tau$		(m)	
	Mass : m/e =	408		( $M^+$ )	
		377		( $M^+ - \text{CH}_3\text{O}$ )	
		349		( $M^+ - \text{CH}_3 \overset{\text{OC-}}{\underset{\text{O}}{\parallel}}$ )	

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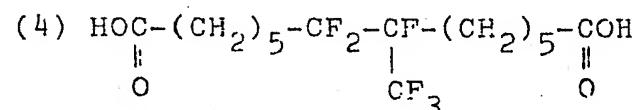


Boiling point: 155-156°C/0.8 - 1 mmHg



Elementary analysis

	C	H	O	F
Calcd. (%)	40.60	4.51	12.03	42.86
Found (%)	40.81	4.52	11.81	42.68
IR:	1020 $\text{cm}^{-1}$		(C-F)	
	1100 - 1230 $\text{cm}^{-1}$		(C-F)	
10	1710 $\text{cm}^{-1}$		(C=O)	
$^1\text{H}$ NMR (CCl <sub>4</sub> ):	11.66 $\tau$		(COOH)	
	0.7 - 2.9 $\tau$		(CH <sub>2</sub> and CFH)	



Elementary analysis

	C	H	O	F
Calcd. (%)	47.37	5.79	16.84	30.00
Found (%)	47.58	5.70	16.92	29.80
IR:	1020 $\text{cm}^{-1}$		(C-F)	
	1100 - 1230 $\text{cm}^{-1}$		(C-F)	
20	1710 $\text{cm}^{-1}$		(C=O)	
$^1\text{H}$ NMR (CCl <sub>4</sub> ):	10.63 $\tau$		(COOH)	
	0.5 - 2.7 $\tau$		(CH <sub>2</sub> and CFH)	

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Example 2

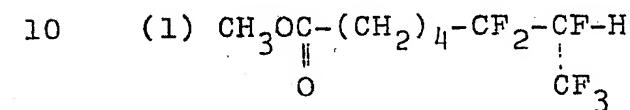
A 150 ml quantity of absolute methanol is placed in a reactor cooled to 0 to 5°C and equipped with a stirrer. Thereto are added 42.0 g of cyclopentanone, 5 3 g of concentrated sulfuric acid and 48.6 g of a 35% aqueous solution of hydrogen peroxide. The mixture is stirred for 60 minutes to yield methoxycyclopentanone peroxide. Absolute methanol (150 ml) is placed in another reactor cooled to -40°C, equipped with a 10 stirrer and filled with a nitrogen gas. Thereto are added 150 g of ferrous sulfate heptahydrate ( $\text{FeSO}_4 \cdot 7\text{H}_2\text{O}$ ) and 80 g of liquefied hexafluoropropene to prepare a suspension. The methoxycyclopentanone peroxide solution is dropwise added by degrees to the suspension with 15 stirring while maintaining the temperature of the suspension at  $-40 \pm 2^\circ\text{C}$ . After the addition thereof, the mixture is stirred for 2 hours and the resulting mixture is left to stand until it is brought to room temperature. The unreacted hexafluoropropene (about 20 62 ml) is recovered by distillation. The reaction mixture is treated at an increased temperature of up to 40°C and a reduced pressure of 10 to 15 mmHg to recover methanol. Then the reaction mixture is separated into 25 an organic layer (upper layer) and an aqueous layer (lower layer). The organic layer (63.5 g) is washed with acid

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and water and dried. A fraction (light yellow to colorless) is obtained by distillation at 40 to 180°C/1 mmHg. The fraction is a mixture of long-chain monobasic acid, long-chain ester of monobasic acid, long-chain dibasic acid and long-chain ester of dibasic acid all of which contain fluorine. The mixture is separated into each compound by rectification.

Analyses by chromatography, mass spectrometry, etc. confirm the production of the following compounds.



Elementary analysis (for  $\text{C}_9\text{H}_{12}\text{O}_2\text{F}_6$ )

	C	H	O	F
Calcd. (%)	40.60	4.51	12.03	42.86
Found (%)	40.33	4.81	12.16	42.70

15 IR:  $1020 \text{ cm}^{-1}$  (C-F)

$1100 - 1230 \text{ cm}^{-1}$  (C-F)

$1750 \text{ cm}^{-1}$  (C=O)

$^1\text{H NMR}$  ( $\text{CCl}_4$ ):  $3.62\tau$  (s,  $\text{CH}_3\text{O}$ )

$0.80 - 2.50\tau$  ( $\text{CH}_2$  and CFH)

20 Mass : m/e =  $266 (\text{M}^+)$

$235 (\text{M}^+ - \text{CH}_3\text{O})$

$207 (\text{M}^+ - \text{CH}_3\text{OC}(\text{CH}_2)_4\text{CF}_2)$

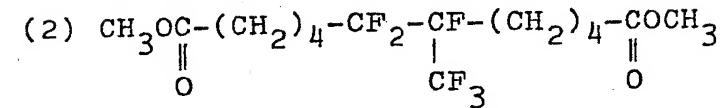
$74 (\text{CH}_2\text{COCH}_3)$

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Boiling point: 58°C/1 mmHg

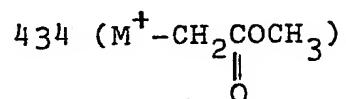
Yield: 44.5 g



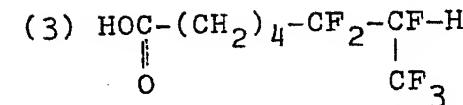
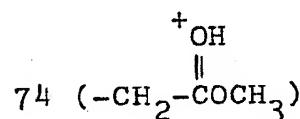
Elementary analysis

5

	C	H	O	F
Calcd. (%)	47.37	5.79	16.84	30.00
Found (%)	47.21	5.78	16.71	30.30

Mass: m/e = 476 ( $\text{M}^+$ )

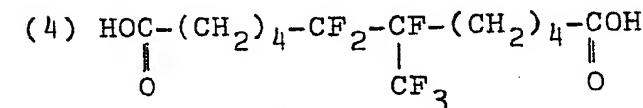
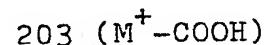
10



Elementary analysis

15

	C	H	O	F
Calcd. (%)	38.10	3.97	12.70	45.24
Found (%)	37.92	4.11	12.80	45.17

Mass: m/e = 252 ( $\text{M}^+$ )

Elementary analysis

20

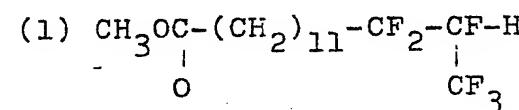
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	C	H	O	F
Calcd. (%)	44.32	5.11	18.18	32.39
Found (%)	44.03	5.11	18.27	32.59

Example 3

5        A 1000 ml quantity of absolute methanol is placed in a reactor equipped with a stirrer and 186 g of 98% cyclododecanone is dissolved thereinto in 20 minutes while maintaining the temperature at 20°C. To the solution are added 10 g of concentrated sulfuric acid and 144.5 g of a 35% aqueous solution of hydrogen peroxide. The mixture is stirred for 60 minutes to yield methoxycyclododecanone peroxide. Absolute methanol (750 ml) is placed in another reactor equipped with a stirrer and filled with a nitrogen gas while retaining 10 the temperature at 20°C. Thereto is added 325 g of 15 ferrous sulfate heptahydrate ( $\text{FeSO}_4 \cdot 7\text{H}_2\text{O}$ ) to prepare a suspension. The methoxycyclododecanone peroxide solution is dropwise added by degrees to the suspension with stirring while maintaining the temperature of the 20 suspension at  $20 \pm 2^\circ\text{C}$  and forcing thereinto a hexa-fluoropropene gas. After the addition of the solution, the gas is continuously forced into the mixture for 2 hours with stirring until the temperature is elevated to 25 35 to 40°C. Thereafter the mixture is further stirred for 4 hours without blowing the gas thereinto. The

mixture is left to stand until it is brought to room temperature, and a methanol is recovered at an increased temperature (up to 40°C) and reduced pressure (10 to 15 mmHg). The reaction mixture is separated into an organic 5 layer (upper layer) and an aqueous layer (lower layer). Since the separation is incomplete, 10 g of 50% dilute sulfuric acid and 50 g of ether are added to the mixture and the resulting admixture is fully stirred. Then the upper layer (ether layer) is decanted, washed with water 10 and dried, followed by the distillation of the ether to prepare a mixture of long-chain monobasic acid, long-chain ester of monobasic acid, long-chain dibasic acid and long-chain ester of dibasic acid all of which contain fluorine. The mixture is found to predominantly 15 comprise the following compounds stated below in (1) and (2).

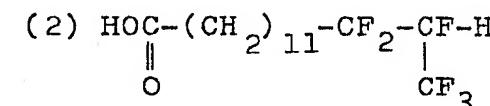


Elementary analysis (for  $\text{C}_{16}\text{H}_{26}\text{O}_2\text{F}_6$ )

		C	H	O	F
20	Calcd. (%)	52.75	7.14	8.79	31.32
	Found (%)	53.00	7.14	8.60	31.26
	IR:	1020 $\text{cm}^{-1}$		1100 - 1230 $\text{cm}^{-1}$	(C-F)
				1740	(C=O)

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Mass: m/e = 364 ( $M^+$ )  
                   333 ( $M^+ - CH_3O$ )  
                   289 ( $M^+ - CH_3OC-$ )



5 Elementary analysis (for  $C_{15}H_{24}O_2F_6$ )

	C	H	O	F
Calcd. (%)	51.43	6.86	9.14	32.57
Found (%)	51.20	6.85	9.04	32.91

Mass: m/e = 236 ( $M^+$ )

10 187 ( $M^+ - COOH$ )

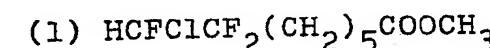
### Example 4

A 27 ml quantity of absolute methanol is placed in a reactor cooled to 0 to 5°C and equipped with a stirrer. Thereto are added 4.6 g (4.7 m mole) of cyclohexanone, 0.16 g of concentrated sulfuric acid and 4.6 g of a 35% aqueous solution of hydrogen peroxide. Then the mixture is stirred for 60 minutes to yield methoxycyclohexanone peroxide.

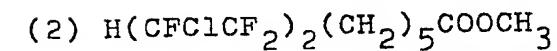
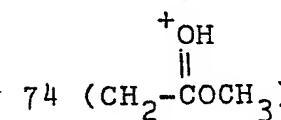
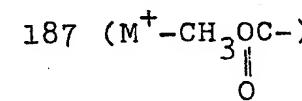
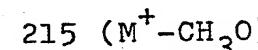
Absolute methanol (27 ml) is placed in another reactor equipped with a stirrer and filled with a nitrogen gas and is cooled to -40°C. Thereto are added 13 g of ferrous sulfate heptahydrate ( $\text{FeSO}_4 \cdot 7\text{H}_2\text{O}$ ) and 8.2 g of liquefied trifluorochloroethylene. The methoxycyclo-

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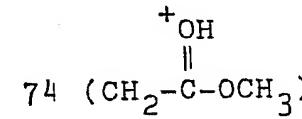
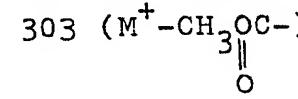
hexanone peroxide solution is dropwise added by degrees to the mixture while maintaining the temperature of the suspension at  $-40^{\circ}\text{C}$ . After the addition of the solution, the stirring is continued at  $-40^{\circ}\text{C}$  for 2 hours and the reaction mixture is left to stand until it is brought to room temperature. The unreacted trifluorochloroethylene and methanol are distilled off and the organic layer is extracted with ether. The extract is washed with acid and water, and dried. Thereafter the ether is distilled off to give 5.0 g of a light yellow oily crude product. Analyses by gas chromatography, mass spectrometry, etc. confirm the production of the following compounds.



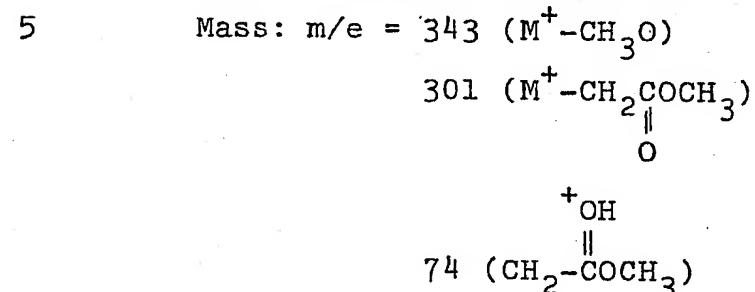
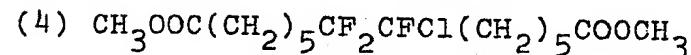
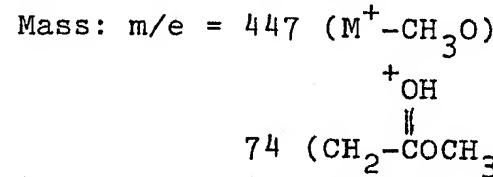
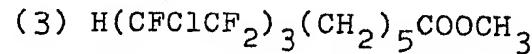
Mass: m/e = 246 ( $\text{M}^+$ )



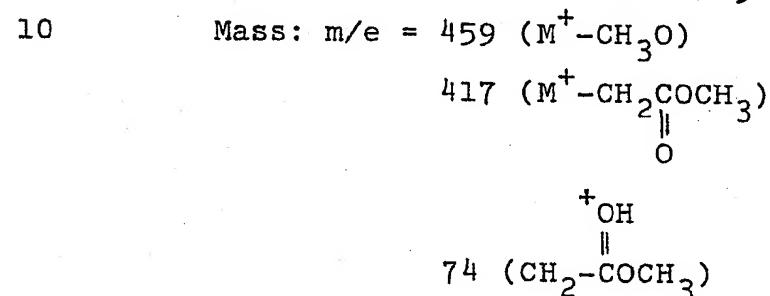
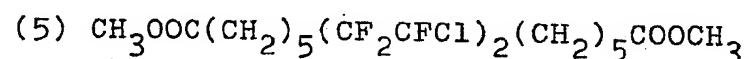
Mass: m/e = 331 ( $\text{M}^+ - \text{CH}_3\text{O}$ )



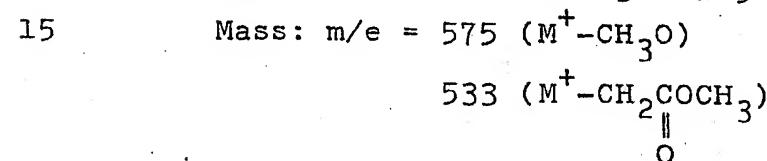
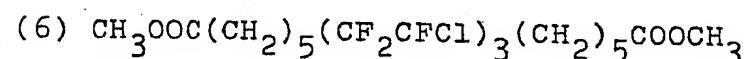
- 27 -



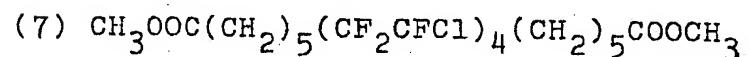
Yield: 0.5 g



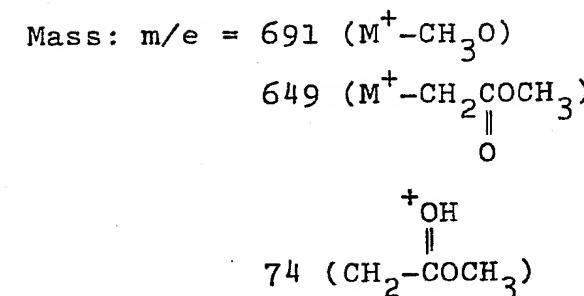
Yield: 0.8 g



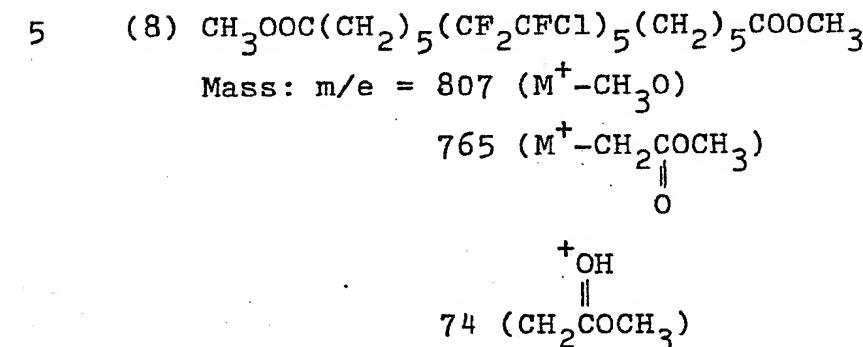
Yield: 1.5 g



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Yield: 1.0 g



(9) Product having a high boiling point

10 Example 5

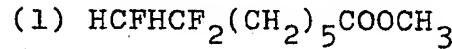
A 9 ml quantity of absolute methanol is placed in a reactor cooled to 0 to 5°C and equipped with a stirrer. Thereto are added 1.47 g of cyclohexanone, 0.05 g of concentrated sulfuric acid and 1.47 g of a 35% aqueous 15 solution of hydrogen peroxide. The mixture is stirred for 60 minutes to yield methoxycyclohexanone peroxide.

Absolute methanol (9 ml) is placed in another reactor equipped with a stirrer and filled with a nitrogen gas and is cooled to -65°C. Thereto are added 4.2 g of 20 ferrous sulfate heptahydrate ( $\text{FeSO}_4 \cdot 7\text{H}_2\text{O}$ ) and 2.5 g of liquefied trifluoroethylene. The methoxycyclohexanone

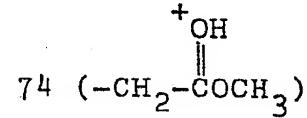
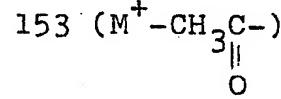
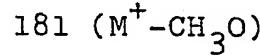
peroxide solution is dropwise added by degrees to the mixture at -68 to -65°C with stirring. After the addition of the solution, the resulting admixture is stirred for 2 hours at -68°C to -65°C and then further for 2 hours at -40°C. Thereafter the reaction mixture is left to stand until it is brought to room temperature.

The unreacted trifluoroethylene and methanol are distilled off and the organic layer is extracted with ether. The extract is washed with acid and water, and dried. Then the ether is distilled away to give 1.28 g of a light yellow oily crude product.

Analyses by gas chromatography, mass spectrometry, etc. confirm the production of the following compounds.



15 Mass: m/e = 212 ( $\text{M}^+$ )

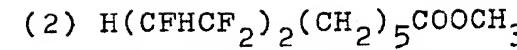


19 FNMR ( $\text{CH}_3\text{COCH}_3$ ): ( $\text{CF}_3\text{COOH}$ ) 33.0 ( $\text{CF}_2$ )

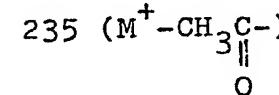
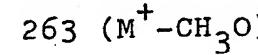
20 44.0 ( $\text{CF}_2\text{H}$ )

119.5 (CF)

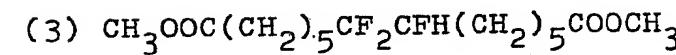
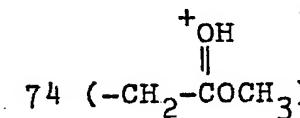
- 30 -



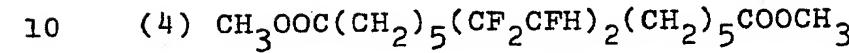
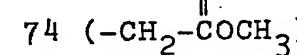
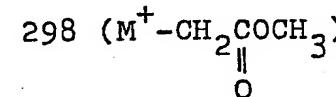
Mass: m/e = 294 ( $\text{M}^+$ )



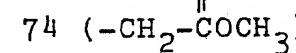
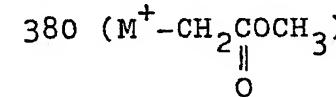
5



Mass: m/e = 340 ( $\text{M}^+$ )



Mass: m/e = 422 ( $\text{M}^+$ )



(5) Products having a high boiling point

15

Example 6

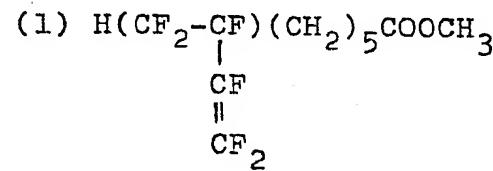
A 9 ml quantity of absolute methanol is placed in a reactor cooled to 0 to 5°C and equipped with a stirrer. Thereto is added 1.47 g of cyclohexanone and the mixture is stirred for 60 minutes to yield methoxy-

cyclohexanone peroxide.

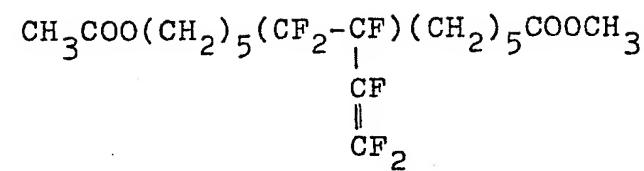
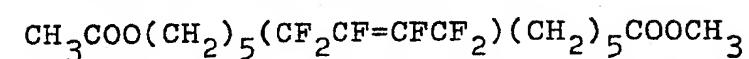
Absolute methanol (9 ml) is placed in another reactor equipped with a stirrer and filled with a nitrogen gas and is cooled to -30°C. Thereto are added 4.2 g of 5 ferrous sulfate heptahydrate ( $\text{FeSO}_4 \cdot 7\text{H}_2\text{O}$ ) and 2.4 g of liquefied hexafluorobutadiene. The methoxycyclohexanone peroxide solution is dropwise added by degrees to the mixture at -30°C with stirring. After continued stirring for 2 hours at -30°C subsequent to the addition of the 10 solution, the reaction mixture is left to stand until it is brought to room temperature.

The unreacted hexafluorobutadiene and methanol are distilled off and then the organic layer is extracted with ether. The extract is washed with acid and water, 15 and dried. Subsequently the ether is distilled away to give 1.5 g of a white wax-like crude product.

Analyses by gas chromatography, mass spectrometry, etc. confirm that a carboxylic acid containing hexafluorobutadiene as fluorine-containing functional groups and 20 the methyl ester thereof are present in the crude product. The crude product is found to predominantly comprise the following compounds.



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(2)  $\text{H}(\text{CF}_2\text{CF}=\text{CFCF}_2)(\text{CH}_2)_5\text{COOCH}_3$ 

(3) Products having a high boiling point

## CLAIMS:

1. A fluorine-containing compound of the formula



5 wherein A is  $\{CF_2CFX\}_m$ ,  $\{CF_2CF\}_m$ ,  $\{CF_2CF\}_m$  or



$\{CF_2CF=CFCF_2\}_m$ ;  $R_1$  and  $R_2$  are each hydrogen or  $R_3OOC(CH_2)_n$ , while  $R_1$  and  $R_2$  may not be hydrogen at the same time; X is hydrogen or halogen;  $R_3$  is hydrogen or a lower alkyl group; m is an integer of 1 to 5; and n is an integer of 10 4 to 11.

2. A fluorine-containing compound as defined in claim 1 in which m is an integer of 1 to 3.

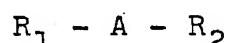
3. A fluorine-containing compound as defined in claim 1 in which A is  $-CF_2CF-$ .



15 4. A fluorine-containing compound as defined in claim 1 in which n is 5.

5. A fluorine-containing compound as defined in claim 1 in which  $R_3$  is hydrogen or a methyl group.

20 6. A process for preparing a fluorine-containing compound of the formula



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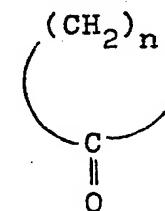
wherein A is  $\langle CF_2CFX \rangle_m$ ,  $\langle CF_2CF \rangle_m$ ,  $\langle CF_2CF \rangle_m$  or

$$\begin{array}{c} CF_3 \\ | \\ CF \\ || \\ CF_2 \end{array}$$

$\langle CF_2CF=CFCF_2 \rangle_m$ ;  $R_1$  and  $R_2$  are each hydrogen or  $R_3OOC(CH_2)_n$ , while  $R_1$  and  $R_2$  may not be hydrogen at the same time; X is hydrogen or halogen;  $R_3$  is hydrogen or a lower alkyl group; m is an integer of 1 to 5; and n is an integer of 4 to 11 which comprises;

5 (i) the first step of reacting a hydrogen peroxide

with a cyclic ketone of the formula



10 (wherein n is an integer of 4 to 11) to give a cyclic ketone peroxide, and

(ii) the second step of reacting the cyclic ketone peroxide thus obtained with a fluorine-containing olefin of the formula

15  $CF_2=CFR_4$

wherein  $R_4$  is hydrogen, halogen, a trifluoromethyl group or the group  $CF_2=CF-$  in the presence of a ferrous salt to give a fluorine-containing carboxylic acid or an ester thereof.

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## EUROPEAN SEARCH REPORT

Application number

EP 80301986.8

DOCUMENTS CONSIDERED TO BE RELEVANT			CLASSIFICATION OF THE APPLICATION (Int. Cl. 5)	
Category	Citation of document with indication, where appropriate, of relevant passages	Relevant to claim	TECHNICAL FIELDS SEARCHED (Int.Cl. 5)	
	<p><u>AT - B - 310 134</u> (CIBA-GEIGY AG) + Pages 3,5-9 + --</p> <p><u>DE - B - 1 196 183</u> (MINNESOTA MINING AND MANUFACTURING COMPANY) + Totality + --</p> <p><u>US - A - 3 248 419</u> (MURRAY HAUPT-SCHEIN et al.) + Columns 6-14 + --</p> <p><u>US - A - 2 905 712</u> (JOHN B. BRAUN-WARTH et al.) + Totality + -----</p>	1-5 1-5 1-5	C 07 C 53/21 C 07 C 55/32 C 07 C 57/52 C 07 C 69/63 C 07 C 69/65// (B 01 F 17/00)	
		6	TECHNICAL FIELDS SEARCHED (Int.Cl. 5)	
			C 07 C 53/00 C 07 C 55/00 C 07 C 57/00 C 07 C 69/00	
CATEGORY OF CITED DOCUMENTS			&: member of the same patent family, corresponding document	
X: particularly relevant A: technological background O: non-written disclosure P: Intermediate document T: theory or principle underlying the invention E: conflicting application D: document cited in the application L: citation for other reasons				
X	The present search report has been drawn up for all claims			
Place of search	Date of compilation of the search	Examiner		
VIENNA	14-08-1980	HOFBAUER		